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1:2. The answer is reasonable and correctly shows two significant

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the precipitate,  $\text{PbI}_2(\text{s})$ :  $2 \text{ mol PbI}_2 - 2 \text{ mol PbI}_2$   
 $\text{PbI}_2$  mol KI 1.785 1 0 mol KI 1 mol  $\text{PbI}_2$   
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the smaller amount. Amount in moles, n, of  
the precipitate, PbI<sub>2</sub>(s):  $2.2 - 2.2$  PbI<sub>2</sub> - 2.2  
PbI<sub>2</sub> mol KI 1.785 1.0 mol KI 1 mol PbI<sub>2</sub>  
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Amount in moles,  $n$ , of the precipitate, PbI

2(s):  $2 \text{ PbI}_2 - 2 \text{ PbI}_2 - 2 \text{ PbI}_2$  2 mol KI 1.785

1 0 mol KI 1 mol PbI

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