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Trends in Atomic Size Influenced by three factors: 1. Energy Level Higher energy level is further away. 2. Charge on nucleus More charge pulls electrons in closer. 3. Shielding effect Group trends As we go down a group... each atom has another energy level, so the atoms get bigger. Periodic Trends As you go across a period, the radius gets smaller.

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Chemical Periodicity Review Sheet . H/Chemistry . A. Completion . Use this completion exercise to check your knowledge of the terms and your understanding of the . concepts introduced in this chapter. Each blank can be completed with a term, short phrase, or number. The periodic table organizes the elements into vertical groups and horizontal ...

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Mendeleev's model for the period table more powerful than other models

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7. Give the names and chemical symbols for the elements that correspond to these atomic numbers: a. 10 b. 18 c. 36 d. 90 8. List, by number, both the period and group of each of these elements. Symbol Period Group a. beryllium Be b. iron Fe c. lead Pb 9. Which of the following pairs of elements belong to the same period?

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Answer. 1. Representative elements: The Representative Elements are those elements within the first two families (Groups I and II on the far left) and the last six families or groups (on the right) of the Periodic Table. 2. d-transition elements: Set of metallic elements occupying a central block (Groups IVB-VIII, IB, and IIB, or 4-12) are d-transition elements.

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